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called components of the solution. On the basis of number of components a solution of two components is called binary solution.
Solute and Solvent

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1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging.
2. A solution may be classified as solid, liquid or a gaseous solution.
- 3.

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Solution Chemistry Notes. 2. Solutions • Basic definitions –
Solution: homogenous mixture – Solute: substance that is dissolved
– Solvent: substance that contains the solute in a solution • Process
of Dissolving: – Molecules of the solvent “pull apart” the solute

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molecules, and distributes them within the solvent • How to tell the solute from the solvent – The substance that is in the greater amount is the solvent – The substance that keeps its phase is the solvent.

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CBSE Class 12 Chemistry Notes: Solutions – Basics. by Anand Meena. April 22, 2019. in 12th Class, Class Notes. Reading Time: 2min read 0. 4. SHARES. 0. VIEWS. Share on Facebook Share on Twitter. Solution is a homogeneous mixture of two or more chemically non reacting substances whose composition can be varied within certain limits. A solution ...

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Solutions Class 12 Notes Chapter 2. A solution comprises a solute

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and a solvent. It is defined as a homogeneous mixture of two or more than two substances. They can be classified into three types: Solid solutions. Gaseous solutions. Liquid solutions. Molarity, mole fraction, percentages, and molality are the terms used to express the concentration of a solution.

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Revision Notes on Solution: A solution is a homogeneous mixture

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of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute. If both components are in same physical state.

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