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$\text{CuO (s)} + \text{H}_2 \text{(g)} \rightarrow \text{Cu (s)} + \text{H}_2 \text{O (g)}$ Here, the oxidation number of Cu decreases from +2 in CuO to 0 in Cu i.e., CuO is reduced to Cu. Also, the oxidation number of H increases from 0 in H₂ to +1 in H₂O i.e., H₂ is oxidized to H₂O. Hence, this reaction is a redox reaction.

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Redox Reactions is the eighth chapter in the NCERT class 11 chemistry textbook. This chapter is regarded by many as

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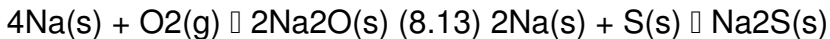
one of the most important chapters in the CBSE class 11 chemistry syllabus, owing to the fact that the entire field of electrochemistry deals with redox reactions.

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the study of the interconversion of electrical and chemical energy during a redox reaction in either voltaic or electrolytic cell.

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redox reactions. 8.2 REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER REACTIONS We have already learnt that the reactions $2\text{Na(s)} + \text{Cl}_2\text{(g)} \rightarrow 2\text{NaCl (s)}$ (8.12)

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(8.14) are redox reactions because in each of these reactions sodium is oxidised due to the addition of either oxygen or more

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Redox reaction involves two half-reactions, one involving loss
of electron or electrons (oxidation) and the other involving
gain of electron or electrons (reduction). ION ELECTRON
METHOD FOR BALANCING REDOX REACTIONS. This

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method involves the following steps: (a) Divide the complete equations into two half reactions-(i) One representing oxidation

Redox Reactions Class 11 Notes Chemistry Chapter 8 ...
Redox Reactions Oxidation is loss of electrons and reduction is gain of electrons. Oxidation and reduction always occur simultaneously, thus the name REDOX reaction. Elements which lose electrons are reducing agents and those which gain electrons are oxidising agents.

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The Class 11 Chemistry Redox Reaction chapter is an essential chapter of the Class 11 syllabus. The chapter has

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high importance in the board exams as well as in the JEE and NEET examinations. So, the students must have adequate knowledge in this chapter. The important topics and sub-topics are given below:

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In the reaction, since H_2O_2 oxidise, O_3 into O_2 thus it behaves as an oxidising agent. the further reaction, in the reaction, Here H_2O_2 reduces Ag_2O into metallic silver $[\text{Ag}]$ (as oxidation number is reducing from +1 to 0). Thus, H_2O_2 behaves as a reducing agent.

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