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the elements are
periodic
functions of
their atomic
numbers.

periodic table.
An arrangement
of the elements
in order of
their atomic
numbers so that
elements with
similar
properties fall

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in the same
column, or
group.

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ANSWER Answer
the following
questions in the
space provided.

1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's

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Original design.
(c) according to
increasing
atomic number.
(d) based on
when they were
discovered.

5 The Periodic Law

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lanthanide.

actinide. alkali
metals. The

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elements with
the atomic
numbers 58-71.
The 14 elements
with the atomic
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group 1 of the
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Describe... 1)

Dalton's model –

solid sphere,

made up of only

1 type of

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matter, no
space. 2)
Schrodinger's
model – based on
probability of
finding an e-;
electrons are
found in
orbitals (look
like electron
clouds), which
are in sublevels
(s,p,d,f,g),
which are in

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energy levels.

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ANSWER Answer

the following

questions in the

space provided.

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1. Chemistry When an electron is added to a neutral atom, energy is (a) always absorbed. (c) either absorbed or released. (b) always released. (d) neither absorbed nor released. 2.

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The Periodic Law

- a. 2 and 4 c. 4
and 5 b. 3 and 4
d. 5 and 6 _____

21. The discovery of the noble gases changed Mendeleev's periodic table by adding a new.
a. period. c.
group. b.

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sublevel block.

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22. In the

Review
modern periodic

Answers
table, elements

are ordered

according to. a.

decreasing

atomic mass. c.

increasing

atomic number.

b. Mendeleev's

original ...

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Quia*

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Chemical Bonding

SECTION 5 SHORT

ANSWER Answer

the following
questions in the
space provided.

1. Identify the
major assumption
of the VSEPR
theory, which is

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used to predict
the shape of
atoms. Pairs of
valence

electrons repel
one another. 2.

In water, two
hydrogen atoms
are bonded to
one oxygen atom.
Why isn't water
a linear ...

6 *Chemical*

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