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photon of light, it absorbs the quantum of energy that the photon carries and moves to a higher energy state. One way of thinking about this higher energy state is to imagine that the electron is now moving faster (it has just been hitting a fast-moving photon).

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Q. Consider the transition of  $n=5$  to  $n=2$  on the Bohr model of the hydrogen atom. This electron transition \_\_\_\_\_ energy.

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Obviously, a positive energy change means that the electron absorbs energy, while a negative energy change implies a release of energy from the electron. Note that the formula is the energy per mole, rather

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Key than that of a single photon. During transition, an electron absorbs/releases energy is in the form of light energy. The energy of the photon

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