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Chapter 14-3

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*Exp-1 To
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fluoride*

*concentration in
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*sample using Ion
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Electrode by

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You will be
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Determination of
concentration of
chloride in an
unknown sample
at the ppm
level. The
chloride ion
selective
electrode you
will use is a
crystalline
solid-state
electrode that
contains a
membrane, as

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shown in the diagram below (Figure 1). The membrane consists of a solid salt of silver sulfide / silver chloride ($\text{Ag}_2\text{S} / \text{AgCl}$). The membrane must be insoluble in the analyte solution and contain the

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Determination of
an analyte ion of
interest.

Of Chloride

Using
Determination of
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Potentiometry -
Chemistry ...

Determination of
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Potentiometry 1.
Purpose This
procedure will
determine the
concentration of

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chloride ion
with a chloride
specific ion
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2. Background
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Potentiometry is
an
electrochemical
method in which
the potential of
an
electrochemical
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Potentiometry is

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method in which

the potential of

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cell is measured

while little to

no current is

passed through

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Determination of
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Purpose

Determination of
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Purpose This
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with a chloride specific ion electrode using potentiometry.

2. Background
Potentiometry is an electrochemical method in which the potential of an electrochemical cell

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Determination of Chloride using Potentiometry 1. Purpose 2.

This procedure will determine the concentration of chloride ion with a chloride specific ion electrode using potentiometry.

2.

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Purpose 2.

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1963

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1984-06-29

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2.1 Total solubilized chloride is determined potentiometrically using a chloride ion-selective electrode (ISE) in conjunction

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with a double-junction reference electrode, or a chloride combination ISE, and a pH meter with an expanded millivolt scale or an ISE meter capable of being calibrated directly in terms of

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chloride
concentration.

METHOD 9212
POTENTIOMETRIC
DETERMINATION OF
CHLORIDE IN ...

The most simple
method for the
determination of
chlorides is to
titrate the milk
directly, using
potassium

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chromate as indicator. With practice, reasonably reproducible results can be obtained, though the results are higher than chloride determinations made after dry ashing of the milk.

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and tributylphos-
phate in
chloroform.

Microchemical
Journal 1990, 41

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Q. Ziling Lu,

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Study of

CI04--selective

electrode based

on a conducting

polymer

polypyrrole. ...

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Using

Precipitation
Titration:

Determination of
Chloride by the
Mohr Method by
Dr. Deniz

Korkmaz

Introduction

Titration is a
process by which

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the determination of the concentration of an unknown substance in solution is determined by adding measured amounts of a standard solution that reacts with the unknown. Then the concentration of

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Determination of Chloride
Using
the unknown can
be calculated
using ...

Precipitation
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Determination of
Chloride by the
...

Besides acid-
base titrations,
the titrimetric
determination of
chloride is one

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of the most frequently used titrimetric methods of analysis. It is employed more or less frequently in practically every laboratory. This Bulletin shows you how to determine chloride in a

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wide range of concentrations using automatic titrators.

Silver nitrate is normally used as titrant (for environmental reasons one ...

Chloride titrations with potentiometric indication

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Using Potentiometry
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electrode that students assemble to determine chloride in natural waters using potentiometric titration with silver nitrate. In our analytical chemistry laboratory, we

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have used this electrode to determine the chloride content in the urban Muddy River in Boston, MA, that is frequently contaminated

Potentiometric
Determination of
Chloride in
Natural Waters

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Determination

7.5.1 Calculate the chloride ion concentration in the original sample, in milligrams per liter, as follows:

$$\text{Chloride (mg/L)} = [(V_1 - V_2) \times N \times 71,000] / S$$

where: V_1 =
Milliliters of

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standard AgNO₃ solution added in titrating the sample prepared in Sec. 7.1. V_2 = Milliliters of standard AgNO₃ solution added in titrating the sample prepared in Sec. 7.3.

METHOD 9253

CHLORIDE

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(TITRIMETRIC,
SILVER NITRATE)
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Determination is calculated using the following equation: $\text{pH} =$

$\log H^+ + 0$

Acidic solutions have a pH value

of less than seven. Solutions

with pH values greater than

seven are described as

basic, or

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alkaline. We use the pH of water as our definition of neutrality. Water is actually a mixture of molecular water (H_2O), and ionised water (H_3O^+ and OH^-). In

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Potentiometry:

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The number of
papers

describing

potentiometric

determination of

chlorides over

the past ten

years is several

times higher

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than for other
halides. On the
other hand, most
of newly
proposed
methods...

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Natural Waters

...

Dispensing
liquid. Hold the

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tip at an angle against the inside wall of the tube/flask.

Press down the control button slowly to the first stop (measuring stroke) and wait until the liquid stops flowing. Press down the control button

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to the second
stop (blow-out)
until the tip is
completely
empty.

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Appendix A:
Operation of
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Adjustable
Pipettes ...
Direct
measurement of
solutions

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containing
sodium chloride
is possible by
reference to a
calibration
based on the
potential
produced by the
electrode pair
as a function of
 pNa^+ defined as
 $-\log_{10}$ sodium
ion
concentration.

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For the measurement of the sodium content of solutions of sodium salts of weak acids and mixed solutions of electrolytes and dextrose it is necessary to use a calibration carried out in a

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buffer system
(0.5M
triethanolamine
+ hydrochloric
acid to pH 7)
and to dilute
the ...

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