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And Salts | Class 10 Chemistry

9th std Science 5.Acids, Bases
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value of important Substance | Gk
for railway ALP, Group D 2018

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Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction. Examples:
 $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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Two solutions A and B have pH values of 5 and 8 respectively. Which solution will be basic in nature ...

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Ans:- Salts are classified into four

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categories on the basis of the strength of acid and base. they are: Salt of strong acid and strong base.e.g NaCl, KCl; Salt of strong acid and weak base. e.g. NH_4Cl , NH_4I ; Salt of weak acid and strong base. e.g. CH_3COONa , $\text{COO} - \text{Na} +$ Salt of weak acid and

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weak base. e.g. $\text{CH}_3\text{COO}^- \text{NH}_4^+$
+ How is common salt found in nature?

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Salts & Chemical Analysis; 9. The
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2. Bases react with . acids to
produce salts and water; water to
produce acids and salts; salts to

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produce acids and water; neither acids, salts, nor water

~~Acids Bases and Salts Multiple
Choice Questions and Answers~~

The salts of strong acids and weak bases give acidic solution having pH less than 7. Example,

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NH_4Cl , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7. Example, Na_2CO_3 , Sodium Carbonate will have pH more than 7. Question 7.

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Answer A reaction in which an acid and base react with each other to give a salt and water is termed as neutralization reaction. For Example: (i) $\text{NaOH} +$

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$\text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (ii) $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ 15. Give two important uses of washing soda and baking soda. Answer Two important uses of washing soda are:

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~~Science: Ch 2 Acids, Bases ...~~

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical

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formula is NaCl . It may be white or light blue or yellow depending upon impurities present in it.

~~Acids Bases and Salts Chapter
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Pattern. phenylamine aniline as
an amine chemguide.

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Taking Answers~~

(A) inorganic acid (B) organic acid
(C) soft organic acid (D) strong
inorganic acid. Answer. Answer:

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(a) 5. When milk of magnesia reacts with acetic acid it produces (A) basic salts (B) acidic salts (C) neutral salt (D) complex salt.

Answer. Answer: (c) 6. Which of the following phenomena will occur when a small amount of acid is added to water ...

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Turmeric is a natural indicator. On

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adding its paste to acid and base separately, which colours would be observed. (a) Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid and yellow in base (d) Red in acid and blue in base. Answer.

Answer: (b) Yellow in acid and red

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subject - acids, alkalis, bases and
salts. In it we find out the

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difference between alkalis and bases, how acids and alkalis react with one another, how neutralisation can produce salts and the relevance of pH values.

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Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

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Acid + Base \rightarrow Salt + Water.

Question 3: Rain becomes acidic due to. carbon dioxide. sulphur dioxide. nitrogen dioxide. all of the above. Correct Option is : 4.

Solution : Rain becomes acidic

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because of carbon dioxide, sulphur dioxide and nitrogen dioxide. They dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid ...

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